

Glue this side  
down into your  
science notebook.

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# Lewis Structure Notes

## What is a Lewis Structure?

A Lewis Structure is a simplified way to represent an atom and its outer electrons. It has two parts to it: the atomic symbol for the element and dots representing the electrons surrounding it.

## Getting Started:

1. Find your \_\_\_\_\_ on the periodic table.
2. Determine the number of \_\_\_\_\_ electrons.
3. This is how many electrons you will \_\_\_\_\_.

## Review:

- On the periodic table, each column is called a \_\_\_\_\_.
- Each element in a \_\_\_\_\_ has the same number of \_\_\_\_\_ in their outer orbital, or \_\_\_\_\_.
- The electrons in the outer shell are known as \_\_\_\_\_ electrons.

## Drawing the Lewis Structure:

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1. Write the \_\_\_\_\_ for the element.
2. How many \_\_\_\_\_ electrons?
3. Start on the \_\_\_\_\_ and draw the electrons counter-clockwise around the element \_\_\_\_\_.
4. Check your work— did you draw the correct number of electrons?

# Photocopy Lewis Structures Handout Inside Booklet

<http://www.middleschoolscience.com/lewis.pdf>